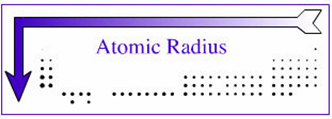
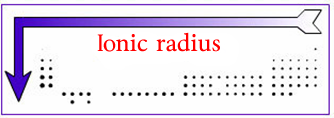
**Describing & Explaining trends in Atomic & Ionic radii (Level 3) examiners tips: Read these please!**

• draw sketches and contrast sizes of atoms and ions sensibly

• see below graphs that summarise the trends in atomic and ionic radii

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#### • note that a positive ion has a smaller radii than its atom

**Also…”don’t be daft”**

#### a comparison and contrast means that different elements/atoms/ions must be discussed in DETAIL, not just the one element/atom/ion

#### protons are never EVER, EVER!!!! lost when an atom forms an ion

#### if describing ionic radii, nuclear attraction on the outer shell electrons is NOT diluted - that's absolute nonsense!

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